

A COMPARATIVE STUDY ON THE ADSORPTION PROPERTIES OF HEAVY METAL Cr IN LAKE SEDIMENT AND SOIL

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Abstract. (1) Background: In order to study chromium (Cr) adsorption characteristics of sediment and shore soil in Xinlicheng Reservoir, Changchun City, Jilin Province, and to discuss the contribution of pH and organic matter adsorption, to explore the accumulation law of Cr in sediment and soil. (2) Methods: OECD Guideline106 equilibrium adsorption method was used in this study. The adsorption kinetics data of Cr in sediment and black soil were fitted by Elovich equation and parabolic diffusion equation. The Freundlich and D-R models were used to fit the adsorption thermodynamic data of Cr. (3) Results: the isothermal adsorption curve of Cr by sediment and soil was "S" type, and K_f sediment $>K_f$ soil. Thermodynamic parameters $\Delta G < 0$, $\Delta H > 0$, to demonstrate the adsorption of Cr was a spontaneous endothermic process, high temperature was beneficial to the spontaneity of adsorption; Within the pH range of the test, higher and lower pH values were not conducive to the adsorption of Cr by sediment and soil. With the increase of pH, the adsorption capacity of Cr by soil and sediment first increased and then decreased. The adsorption amount of Cr in sediment and soil was the smallest at the pH 3, and the adsorption amount was the largest at the pH 7. However, the difference between adsorption of Cr in the sediment under different pH was not significant ($p > 0.05$), and the adsorption of Cr in the soil under different pH was significant ($p < 0.05$). Compared with the soil and sediment without organic matter removal, the adsorption capacity of removing organic matter was significantly reduced ($p < 0.05$), and the contribution rates of organic matter in sediment and soil to Cr were 50.38% and 48.40%, respectively. Consequently, the content of organic matter played a key role in the adsorption of Cr in soil and sediment.

Keywords: chromium, sediment and soil, adsorption characteristics, pH, organic matter content

Introduction

Heavy metals are non-biodegradable accumulative pollutants. The pollution problem is of global concern due to the toxic, persistent and non-degradable nature of excessive heavy metals in the environment. Chromium (Cr), a common heavy metal, is characterized by bioaccumulation, high toxicity, carcinogenicity and mutagenicity, and can cause human cancer, severe diarrhea and kidney damage even at low concentrations (Song et al., 2019). Chromium is among the top hazardous substances priority of the Superfund (Jiang et al., 2018).

Heavy metals in soil mainly come from chemical fertilizers, pesticides, industrial wastewater and irrigation sewage and soil parent material. Chromium (VI) has strong

fluidity, solubility and durability, which leads to polluted water, soil and sediment. Especially during heavy rain, pollutants enter the sediment in the riparian zone along with a large amount of soil, mud and sand, and sediment become the carrier, destination and storage place for various substances to migrate and transform in the water environment (Pan and Ding, 2011; Drinčić et al., 2018; Zhang et al., 2019). Through adsorption or chemical and physical combination, soil and sediment tend to accumulate heavy metals. They are both carriers and reservoirs for heavy metals, which play an important role in the migration of heavy metals.

In recent years, soil heavy metal pollution has received considerable attention by many researchers, and related studies have been published. These researches have shown that adsorption is an important process to control the presence of heavy metals in solid phase, including soil and sediment. Adsorption and desorption have been demonstrated closely related to soil properties, including pH, organic matter, cation exchange capacity, and clay minerals, etc. A large number of studies have shown that pH is a main contributing factor. Zhang et al. (2017) studied the migration rule of Cr(VI) in the soil of vegetable base, and found that the initial concentration of heavy metal Cr and environmental pH value exerted an influence on the adsorption process of Cr(VI). Liu et al. (2018) proposed that the adsorption capacity of the adsorbent is closely related to pH, which may change the existing form of Cr, thus changing the leaching rate of heavy metals in the soils. Dong et al. (2009) studied the adsorption of Cr(VI) in moderately polluted farmland soil in northeast China and the effect of pH value on the adsorption, and found that the adsorption capacity of Cr(VI) in soil decreased against the increase of pH, and the migration rate of Cr(VI) in soil and groundwater increased, which caused greater harm to the environment. Fonseca et al. (2009) studied the adsorption of Cr by acidic soils in Portugal and found that the mobility of Cr(VI) was low in acidic soils. In addition to soil pH, soil organic matter is also one of the important properties that affect the availability of heavy metals. Major contributor to the ability of exchangeable forms of organic matter to retain heavy metals. Zeng et al. (2011) found that Cr content in soil and Cr content in rice tissue were negatively correlated with soil pH value, but positively correlated with organic matter content. Vega et al. (2006) found that the main soil components affecting the adsorption of one or more heavy metals were organic matter, Fe oxide, sidrite and chlorite content, and CEC. In order to probe into the heavy metal pollution of the soil of the Luona River deposit in Tibet, and the surrounding sediment, Qiao et al. (2020) conducted a study and found that Cr pollution in the soil mainly comes from river water and sediment polluted by heavy metals.

The present study focuses on the migration of heavy metal Cr in the sediment of Xinlicheng Reservoir, and the soil at the entrance of the reservoir, Changchun city, Jilin Province, China, and identifies the value of soil pH and organic matter content as influencing factors so that scientific basis is provided for the future restoration of heavy metal contaminated soil around lakes and reservoirs.

Materials and methods

Samples for test

The basin of Xinlicheng Reservoir is slightly rectangular in shape with an average width of 20.7 km. Chromium pollution sources in the water of Xinlicheng Reservoir mainly come from sewage discharge from surrounding factories and Cr migration

caused by rainfall. The soil samples were collected from the river inlet, and the surface layer of 0-20 cm soil was collected by quartering method. The soil type was black soil, the land is agricultural cultivated land, the crop is maize, and NPK compound fertilizer is applied. The sediment is collected at surface layer (0-10 cm) by grab dredger, that is, sampling points are equidistant from the upstream of the reservoir (5 km, 10 km, 15 km away from the reservoir), and the sampling point is 2 m away from the shore, repeated 3 times for each collection. After the samples collected at the three sampling points were mixed and naturally air-dried indoors, the impurities such as animal and plant residues and rocks were removed, and then the samples were ground through a 10 and 100-mesh sieve for storage. The basic physical and chemical properties of the tested soil samples are shown in *Table 1*.

Table 1. Physicochemical properties of test soil

Sample	Unit weight (g·m ⁻³)	OM/ (g/kg)	T _N / %	T _P / %	Cr /mg·kg ⁻¹	Mechanical composition%			
						Sand	Coarse silt	Fine silt	Clay
Sediment	1.79	41.23	0.15	0.12	33.73	10.48	33.97	24.12	31.43
Black soil	1.65	34.65	0.12	0.09	26.55	4.84	51.43	22.96	20.77

Test design

Adsorption kinetics

Adsorption kinetics test was used to evaluate the equilibrium time of subsequent adsorption isotherm tests. In the adsorption kinetics test, soil and water ratio 20:1 was used in 50 ml polyethylene centrifuge tube, and 1.00 g soil sample and 20 ml Cr solution with a concentration of 200 mg·L⁻¹ were added, respectively. They oscillate for 1, 5, 10, 20, 30, 60, 120, 240, 360, 480, 720, 1440 min at the temperature of 25°C. After centrifugation at 4000 r·min⁻¹ for 10 min, the concentration of Cr in the supernatant was determined by filtration with 0.45 µm microporous membrane (Dali-Youcef et al., 2006).

Adsorption isotherm

The 1.00 g samples of sediment and soil were added into a 50 mL polyethylene centrifuge tube to study the adsorption isotherm of Cr on sediment and soil. The influence of univariate time on the adsorption of Cr on sediment and black soil was studied when different initial concentrations of Cr were 50, 100, 200, 300, 400 and 500 mg·L⁻¹, respectively. Adsorption kinetics on the basis of constant temperature oscillation at 25°C to obvious equilibrium, the test tube with Cr solution added as blank, centrifugation, filtration, determination and other steps are the same as Adsorption kinetics (Arias et al., 2006).

Influence of temperature on adsorption isotherm

The adsorption of Cr onto sediment and soil at 15°C, 25°C and 35°C was studied by referring to the batch adsorption test. The influence of single variable temperature on the adsorption of Cr onto sediment and black soil at different initial concentrations of 50, 100, 200, 300, 400 and 500 mg·L⁻¹ was determined (Abat et al., 2012).

Influence of pH on adsorption isotherm

The solutions of $0.1 \text{ mol}\cdot\text{L}^{-1}$ HCl and $1.0 \text{ mol}\cdot\text{L}^{-1}$ NaOH were used to adjust the pH values of different Cr solutions to 3, 5, 7, 9 and 11. According to Adsorption isotherm method, adsorption isothermal determination was conducted at 25°C to study the effect of pH change of a single variable on the adsorption of Cr on sediment and black soil.

Influence of organic matter content on adsorption isotherm

The organic matter in sediment and soil was removed by H_2O_2 digestion. The bottom mud and soil samples were weighed after air drying, and soaked with 6% H_2O_2 for 24 h, and after that 30% H_2O_2 100 ml was added for several times. Repeated treatment was made till no bubbles emerged from the samples. The residual H_2O_2 was removed by heating method and finally washed with deionized water.

Organic matter samples were taken respectively (with non-removed organic matter samples as the control), and the adsorption isotherm was determined at 25°C according to Adsorption isotherm method, to study the influence of changes in single variable organic matter on sediment and soil Cr adsorption.

Analysis of adsorption data

Adsorption kinetic model

Elovich equation:

$$q_t = a + b \ln t \quad (\text{Eq.1})$$

Parabolic diffusion equation:

$$S/S_{\max} = a + bt^{1/2} \quad (\text{Eq.2})$$

where, a is the adsorption constant related to the maximum adsorption capacity, and b is the adsorption rate coefficient, which can be interpreted as how fast the reaction rate decreases. The Elovich equation is used to explain the adsorption process of heterogeneous diffusion because the activation energy varies greatly during the reaction.

Isothermal fitting equation (Fan et al., 2012):

The Freundlich equation:

$$q_e = K_F C_t^{1/n} \quad (\text{Nonlinear form}) \quad (\text{Eq.3})$$

D-R equation:

$$\ln q_e = \ln q_m - k \varepsilon^2 \quad (\text{Eq.4})$$

where, q_e is the adsorbed amount of Cr ($\text{mg}\cdot\text{g}^{-1}$); q_m is the maximum adsorption capacity, k is the constant related to the adsorption energy, ε is the adsorption potential of Polanyi, K_F is the constant, n is the affinity value of the constant. All these formula

symbols respectively represent the maximum adsorption capacity of Cr and the adsorption capacity.

Thermodynamic parameter equation (Chen et al., 2011):

$$\Delta G = -RT \ln K_F \quad (\text{Eq.5})$$

$$\Delta G = \Delta H - T\Delta S \quad (\text{Eq.6})$$

Type: ΔG for the change of Gibbs free energy ($\text{kJ}\cdot\text{mol}^{-1}$), $\Delta G < 0$ indicates that reactions can occur spontaneously, $\Delta G > 0$ reaction can't occur spontaneously; ΔH is the standard enthalpy change ($\text{kJ}\cdot\text{mol}^{-1}$), $\Delta H < 0$ indicates the reaction is exothermic, while $\Delta H > 0$ indicates the reaction is endothermic; $T(\text{K})$ is the absolute temperature; R is the ideal gas constant = $8.314 \times 10^{-3} \text{ kJ}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$; ΔS is the standard entropy change ($\text{kJ}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$), $\Delta S < 0$ means chaotic degree is reduced, $\Delta S > 0$ indicates confusion degree is increased; K_F stands for thermodynamic equilibrium constant. When the $1/T$ and $\ln K$ are used to draw horizontal and vertical curves, ΔH and ΔS can be obtained according to the slope and intercept, respectively.

All tests were performed in triplicate.

Results and discussion

Adsorption kinetics of Cr

Figure 1 depicts the adsorption of heavy metal Cr in sediment and soil at 25°C and shows that the whole adsorption process of metal Cr in sediment and soils consist of two stages, the rapid adsorption stage and the slow equilibrium adsorption stage. Within 0-120 min, with the increase of adsorption time, the adsorption amount of Cr by sediment and black soil increases significantly, and the concentration of Cr in solution decreases rapidly. However, within 120-1440 min can be categorized into the slow equilibrium stage. At this stage, the change of Cr concentration in solution gradually tends to be in equilibrium. Kinetic test demonstrates that the adsorption equilibrium time of Cr in sediment and soil is 1440 min, and the maximum adsorption amount of Cr in sediment and black soil is $1665.02 \text{ mg}\cdot\text{kg}^{-1}$ and $1079.04 \text{ mg}\cdot\text{kg}^{-1}$, respectively. In the rapid adsorption stage, the adsorption capacity of Cr by sediment and black soil reaches 79.21% and 78.97% of the total adsorption capacity, respectively. With the passage of adsorption time, the heavy metal Cr in the solution is in full contact with the sediment and the black soil, and gradually occupies the adsorption site in the adsorbent, so the adsorption capacity of the sediment and the black soil gradually increases (Tokunaga et al., 2001; Li et al., 2013). The higher adsorption capacity of Cr in sediment may be due to the high content of organic matter in sediment, which can provide more adsorption points for Cr (Liu and Dai, 2015). In the process of Cr adsorption by sediment and soil, some components in the sediment and soil may form complex with Cr, transfer the Cr adsorbed on the surface of sediment and soil to the interior of particles, and lead to more Cr forming precipitation on the surface of sediment and black soil. When the adsorption time reaches 2 h, the adsorption amount of Cr in the bottom mud and black soil reaches about 80% of the total adsorption amount, and the equilibrium adsorption state can be

reached at 24 h. Thus, it can be seen that the bottom mud and black soil have a strong accumulation capacity for heavy metal Cr.

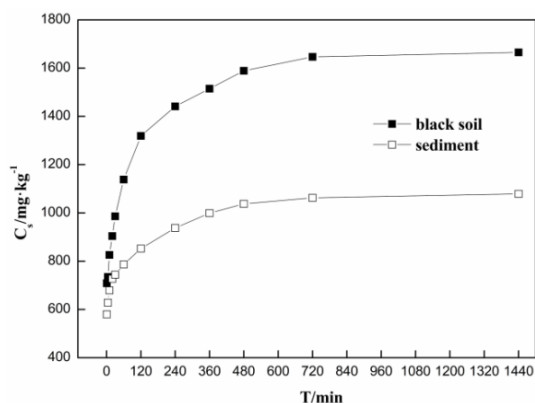


Figure 1. The adsorption kinetics curves of Cr in black soil and sediment

Elovich equation and parabolic diffusion equation were used to fit Cr adsorption kinetic data, so as to understand Cr adsorption mechanism in sediment and soil. Generally, the better the model is, the greater the R^2 value will be. Table 2 shows the fitting parameters of the equation. It can be seen that the adsorption correlation coefficient R^2 of Elovich equation for Cr in sediment and soil was 0.9444 and 0.9569, respectively, which being greater than that of the parabolic diffusion equation (R^2 0.8105 and 0.8485, respectively), indicating that Elovich equation was more suitable for descriptions the Cr adsorption kinetics characteristics in sediment and soil. The higher the constant a value in Elovich equation, the higher the adsorption rate, and therefore the metal Cr is more likely to accumulate in the sediment. Elovich equation is an empirical equation, which is used to describe the adsorption behavior of pollutants on the non-uniform solid adsorption surface and is suitable for the process in which the activation energy changes greatly during the reaction process. The adsorption process involves not only particle diffusion but also multiple forces.

Table 2. Parameters of the kinetics adsorption models for Cr in sediment and black soil

Soil type	Elovich equation			Parabolic diffusion equation			
	a	b	R^2	a	b	s	R^2
Sediment	526.6285	161.1724	0.9444	30.6494	1.0879	26.9135	0.8105
Black soil	512.2667	78.3782	0.9569	122.2293	2.6892	5.3592	0.8485

Note: R stands for correlation coefficient; $p < 0.01$ indicates significant correlation

Adsorption isotherm

The isotherm of heavy metal Cr in sediment and soil at 25°C are introduced in Figure 2 and Figure 3. From the Figures, with the increase of Cr concentration, the adsorption isotherms of Cr on the sediment and soil are on a gradual upward trend of slope, but the adsorption does not reach the saturation state. Within the test concentration range, when the initial concentration of Cr is 500 $\text{mg}\cdot\text{kg}^{-1}$, the equilibrium adsorption capacity of sediment and soil to Cr reaches 8623.43 and 8396.40 $\text{mg}\cdot\text{kg}^{-1}$, respectively. The

adsorption trend of Cr by different adsorbents (sediment and black soil) is roughly the same, but the adsorption amount is significantly different.

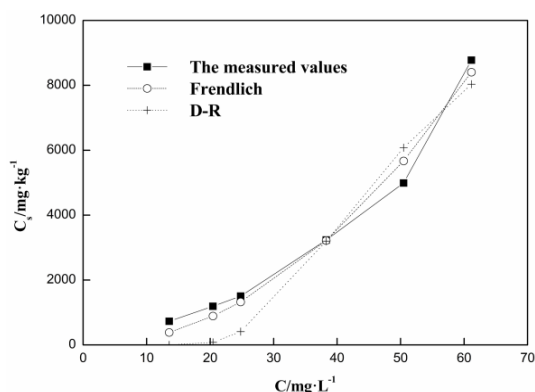


Figure 2. Adsorption isotherms of Cr in sediment

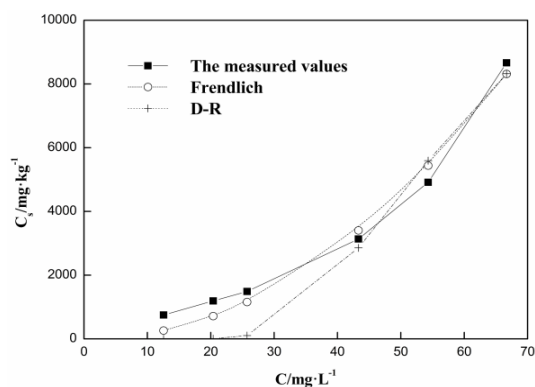


Figure 3. Adsorption isotherms of Cr in black soil

The isothermal adsorption curve of Cr by sediment and soil shows an "S" type. Through the adsorption isotherm data fitting, it is found that the Freundlich (Alemayehu et al., 2011) and D-R model has better data fitting effect than Langmuir model, which can explain the adsorption mechanism of Cr on sediment and soil more reasonably. Table 3 shows the adsorption parameters of Freundlich and D-R. In the Freundlich model fitting, chromium adsorption coefficient R (0.9882 and 0.9834) in sediment and soil is higher than that of D-R model (0.9663 and 0.9675). This is mainly because when the concentration of Cr is less than 500 mg L⁻¹, the adsorption of Cr by sediment and soil does not reach saturation state. The Freundlich equation's adsorption constant K_f represents the adsorption capacity of the adsorbent. K_f sediment is 0.4439 larger than K_f soil, indicating that the sediment is more likely to absorb the metal Cr than the soil.

Table 3. Parameters of the kinetics adsorption models for Cr in sediment and black soil

Soil type	Freundlich model			D-R model			
	K _f	n	r	q _m	b	k	r
Sediment	1.8293	0.4882	0.9882**	14677.4102	0.1789	372.7705	0.9663**
Black soil	1.3854	0.4835	0.9834**	18359.7303	0.0346	581.1108	0.9675**

Note: n=12, r_{0.05}=0.0.576, r_{0.01}=0.708, **Represents significance level (p<0.01)

Adsorption thermodynamics

Temperature is not only an important environmental factor affecting adsorption but also an important condition affecting adsorption membrane properties. The thermodynamic curves of Cr adsorption in sediment and soil are depicted in *Figure 4* and *Figure 5*. It can be seen from the Figures that the adsorption of Cr in sediment and soil are 8623.40, 8776.82, 8911.64 $\text{mg}\cdot\text{kg}^{-1}$ and 8396.40, 8666.46, 8846.86 $\text{mg}\cdot\text{kg}^{-1}$ under the temperatures of 15°C, 25°C and 35°C, respectively. With temperature increasing, the adsorption capacity of Cr in sediment and soil gradually enhances. It indicates that the higher the temperature is, the more favorable it is for Cr adsorbed on the bottom mud and soil surface to diffuse into the particle interior, and conducive to the transformation from thermodynamic unstable compound state to stable compound state.

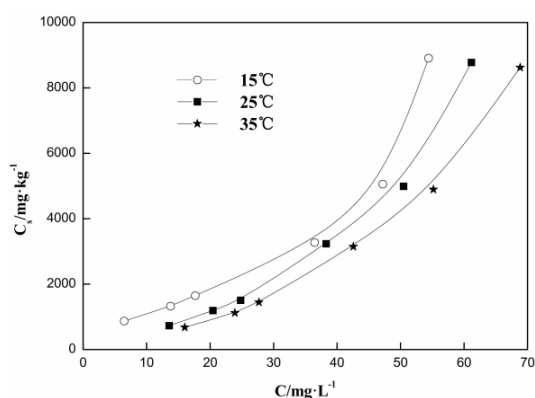


Figure 4. Thermodynamic curves of Cr in sediment

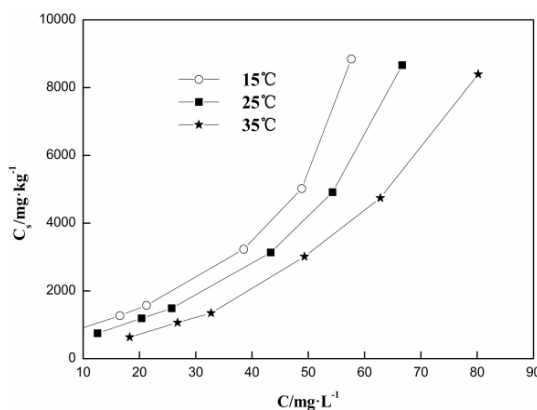


Figure 5. Thermodynamic curves of Cr black soils

Table 4 and *Table 5* show the adsorption thermodynamic parameters of Cr in sediment and soil. It can be seen that the Freundlich model has the best fitting effect, and the Q_m value increases with the increase of temperature, which further demonstrates that high temperature is conducive to the adsorption of Cr in sediment and soil. It can be observed from *Table 5* that within the test temperature range (15, 25 and 35°C), ΔG are less than zero; $\Delta H > 0$ shows that sediment and soil adsorption of Cr process is a

spontaneous endothermic process, and high temperature is beneficial to spontaneous adsorption. The adsorption of heavy metals by adsorbents experiences an endothermic process, and the experimental results are consistent with the trend of (Singh et al., 2017). ΔG and ΔH show that seasonal differences exist in sediment, soil buffer capacity Cr, liquidity and activity of (Lyubchik et al., 2004; Alothman et al., 2013).

Table 4. Adsorption simulation equation parameters for different temperatures

Sample type	Temperature (°C)	Freundlich equation			D-R equation		
		K_f	n	R^2	q_m	k	R^2
Sediment	15	1.430	0.488	0.9902	14279.30	452.5590	0.8990
	25	1.829	0.488	0.9776	14677.41	372.7706	0.8753
	35	3.313	0.513	0.9008	21766.84	470.4402	0.8118
Black soil	15	1.015	0.487	0.9948	14445.38	632.5860	0.9161
	25	1.385	0.483	0.9364	18359.73	581.1089	0.8757
	35	1.756	0.479	0.9382	21138.22	506.8181	0.8435

Table 5. Adsorption thermodynamic parameters of Cr in sediment and soil

Soil type	ΔG			ΔH	ΔS
	15°C	25°C	35°C		
Sediment	-0.8507	-1.9506	-3.0504	30.8503	0.1105
Black soil	-0.0702	-0.6309	-1.3306	20.2406	0.0708

Effect of pH on Cr adsorption

A variety of factors can affect the adsorption of metal Cr by the adsorbent. Figure 6 shows the effect of different pH values on the adsorption of metal Cr. It can be seen that with the increase of pH value, the adsorption capacity of sediment and soil to Cr first increases and then decreases. When the pH value of the experimental environment was 3, chromium has the lowest adsorption capacity in sediment and soil, which was 3090.62 mg·kg⁻¹ and 2256.62 mg·kg⁻¹, respectively. When pH was 7, it reaches the turning point of adsorption, and the maximum adsorption capacity was 3140.02 and 2406.84 mg·kg⁻¹, respectively.

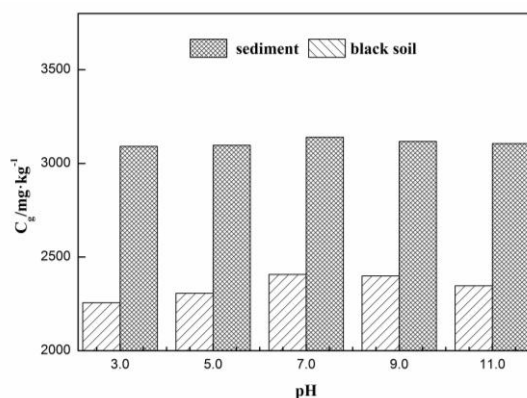


Figure 6. Effect of pH on Cr adsorption in sediment and black soil

When the pH value was within the range of 7-9, the adsorption effect of heavy metal Cr reaches the maximum level in sediment and soils, and the adsorption capacity of Cr was stronger in sediment. When pH was 11, the content of OH⁻ in the solution increases, and the competitive adsorption from H⁺ decreases. At this time, the activity of Cr decreases, thus reducing the migration ability of Cr in soil and sediment. The adsorption amount of Cr in sediment and soil was the smallest when the pH was 3, and the adsorption amount was the largest when the pH was 7. Within the pH range of the test, higher and lower pH values were not conducive to the adsorption of Cr by sediment and soil. With the increase of pH, the adsorption capacity of Cr by soil and sediment first increased and then decreased. The adsorption amount of Cr in sediment and soil was the smallest at the pH 3, and the adsorption amount was the largest at the pH 7. However, the difference between adsorption of Cr under different pH in the sediment was not significant ($p > 0.05$), and the adsorption of Cr under different pH in the soil was significant ($p < 0.05$). This may be due to the higher content of organic matter in the sediment, and higher organic matter makes the sediment have a higher adsorption capacity of Cr, so the impact of organic matter content was greater than the pH value.

The reason for this phenomenon was assumed to be that when pH was 3, the solution contains a large number of H⁺, and the presence of H⁺ produces competitive adsorption with Cr. The pH changes the existing form of Cr in sediment and soil. The lower the pH is, the greater the potential difference of the system will be, thus making it easier for some Cr(VI) to be converted into Cr(III). Cr(III) produces precipitation benefit, reducing the equilibrium concentration of Cr(VI) in solution (Chen et al., 2015). The pH affects H⁺ concentration in sediment and soil. The lower the pH, the higher the H⁺ concentration, the dissociation of hydroxyl group (-OH) on the surface of soil colloidal particles will be affected, and the potential balance of the colloidal group will be destroyed (Liu et al., 2012). With the increase of pH, the content of H⁺ in the solution decreases gradually, competitive adsorption decreases, and OH⁻ increases. The content of calcium carbonate and hydroxyl (-OH) in the colloidal core in the soil was high, which strengthens the adsorption and fixation of Cr by carbonate and oxide in the soil (Zhang et al., 2017). When pH value continues to rise, the increase of OH⁻ concentration will reduce the proton adsorption sites on the surface of adsorbent, enhance the electro negativity of sediment and soil surface, and weaken the adsorption capacity of anionic groups such as HCrO₄⁻, CrO₄²⁻ and Cr₂O₇²⁻, resulting in the decline of Cr adsorption capacity (Wang et al., 2013).

Effects of organic matter on Cr adsorption

As one of the most active solid components in soil chemistry, soil organic matter has a strong adsorption effect on pollutants, thus affecting their migration behavior. When the content of organic matter was high, these organic matter can be complexed with heavy metal ions, thus changing the adsorption property of soil and sediment. The cation exchange capacity in soil and sediment was restricted by oxygen-containing functional groups in humus. Therefore, the humic acid in organic matter can change the cation exchange capacity and thus affect the adsorption capacity of soil and sediment to Cr. *Figure 7* and *Figure 8* show the changes of Cr adsorption before and after removal of organic matter from sediment and soil. It can be seen that before and after the removal of organic matter, when the concentration of Cr added in the solution was at most 500 mg·L⁻¹, the Cr adsorption capacity of sediment reached 8776.82 and 5836.58 mg·kg⁻¹, and the Cr adsorption capacity of soil was 8646.46 and

5826.52 mg·kg⁻¹, respectively, moreover, the difference of adsorption amount for Cr between before and after the removal of organic matter in sediment and soil was significant ($P < 0.05$). According to the formula of contribution rate = $(Q - Q_{\text{removal}}) / Q_{\text{removal}}$, the contribution rate of organic matter in sediment and soil to Cr was 50.38% and 48.40%, respectively. This is related to the organic matter content of the sediment and soil. Organic matter provides large amounts of pollutants adsorption points, and has a strong surface adsorption capacity. The organic matter can reduce hexavalent Cr to trivalent Cr in sediment and soil, and make it into a fixed precipitation state and the organic combination state and lose activity (Agrafioti et al., 2014).

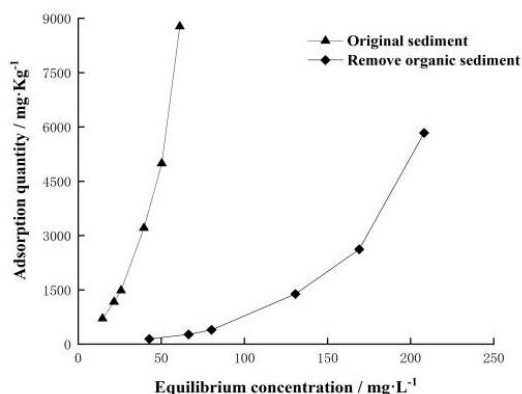


Figure 7. Adsorption isotherm of Cr onto sediment before and after removal of organics

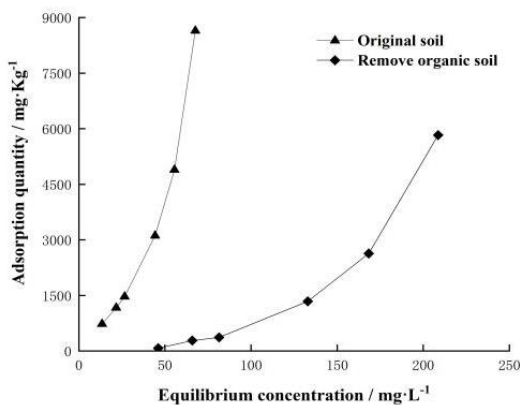


Figure 8. Adsorption isotherm of Cr onto soil before and after removal of organics

Conclusion

(1) Compared with the soil, the lake sediment had a strong adsorption capacity for Cr under the same condition, and the maximum adsorption amount of Cr in sediment and soil was 1,665.02 mg·kg⁻¹ and 1079.04 mg·kg⁻¹, respectively. The isothermal adsorption curve of Cr in sediment and soil showed an "S" type, and the Freundlich equation was better to describe the thermodynamic characteristics of Cr adsorption.

(2) Within the temperature range designed in the experiment, the thermodynamic parameters $G < 0$, $\Delta H > 0$, indicates that sediment and soil adsorption of Cr process is

spontaneous and endothermic, and high temperature is beneficial to the spontaneity of the adsorption.

(3) In the test pH range, higher and lower pH values are not conducive to the adsorption of Cr by sediment and soil. With the increase of pH value, the adsorption capacity of soil and sediment to Cr first increased and then decreased. The adsorption capacity of Cr by sediment and soil is minimum at pH 3 and maximum at pH 7. Therefore, in acidic and alkaline environment, it is not conducive to the fixation of Cr in soil and sediment, which increases the environmental risk of Cr.

(4) The contribution rate of organic matter in sediment and soil to Cr was 50.38% and 48.40%, respectively. After the organic matter is removed, the adsorption capacity was significantly reduced, the activity of Cr increases and resulted in the increase of the environmental risk of Cr.

This article studied the adsorption characteristics of Cr in sediment and soil, and effects of the value of pH and content of organic matter on the adsorption capacity of Cr, in order to seek effective control measures of the Cr environmental risk in soil and sediment through adjusting the environmental conditions such as temperature, the content of organic matter and pH and so on.

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Conflicts of Interests. The authors declare no conflict of interests.

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